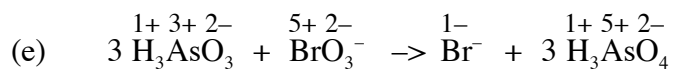


Each K lost $1 e^-$ so K is oxidized & is reducing agent.

Each H gained $1 e^-$ so H in H_2O was reduced & H_2O is oxidizing agent.



Each As lost $2 e^-$ so As in H_3AsO_3 is oxidized & H_3AsO_3 is reducing agent.

Each Br gained $6 e^-$ so Br in BrO_3^- was reduced & BrO_3^- is oxidizing agent.